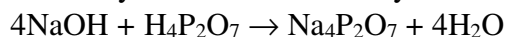
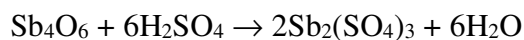


Příklad 2.2.9. Sestavte rovnice následujících dějů:

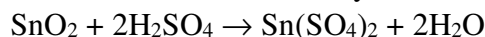
reakce hydroxidu sodného s kyselinou tetrahydrogendifosforečnou



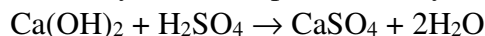
reakce dimerního oxidu antimonitého s kyselinou sírovou



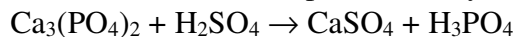
reakce oxidu cíničitého s kyselinou sírovou



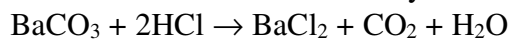
reakce hydroxidu vápenatého s kyselinou sírovou



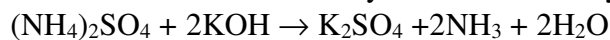
reakce fosforečnanu vápenatého s kyselinou sírovou



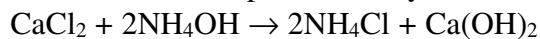
reakce uhličitanu barnatého s kyselinou chlorovodíkovou



reakce síranu amonného s hydroxidem draselným



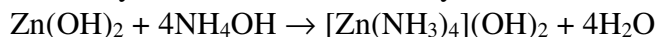
reakce chloridu vápenatého s hydroxidem amonným



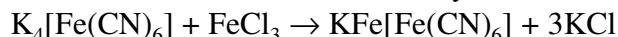
reakce oxidu stříbrného s hydroxidem amonným



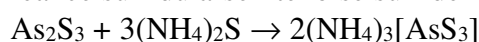
reakce hydroxidu zinečnatého s hydroxidem amonným



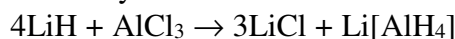
reakce chloridu železitého s hexakynoželeznatanem draselným



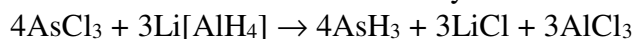
reakce sulfidu arsenitého se sulfidem amonným



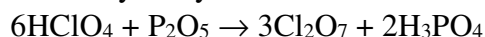
reakce hydridu lithného s chloridem hlinitým



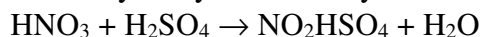
reakce chloridu arsenitého s tetrahydridohlinitanem lithným



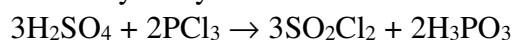
reakce kyseliny chloristé s oxidem fosforečným



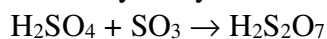
reakce kyseliny dusičné s kyselinou sírovou



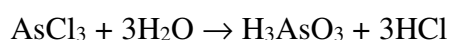
reakce kyseliny sírové s chloridem fosforitým



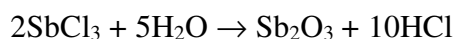
reakce kyseliny sírové s oxidem sírovým



reakce chloridu arsenitého s vodou



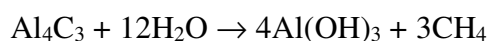
reakce chloridu antimonitého s vodou



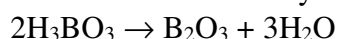
reakce dimerního oxidu fosforečného s vodou



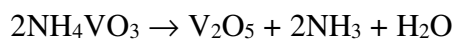
reakce karbidu hlinitého s vodou



termického rozkladu kyseliny trihydrogenborité



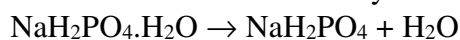
termického rozkladu vanadičnanu amonného



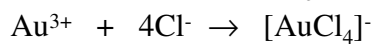
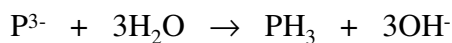
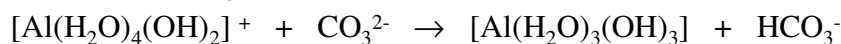
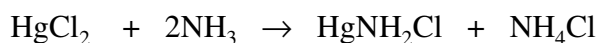
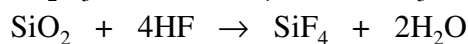
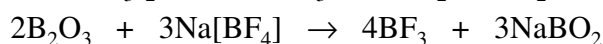
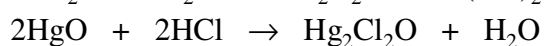
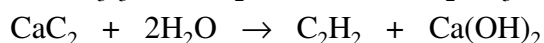
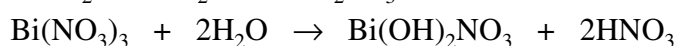
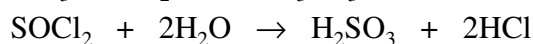
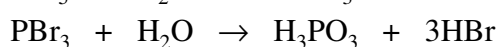
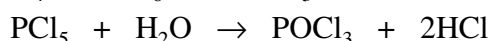
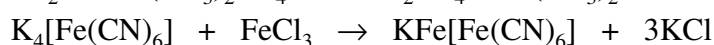
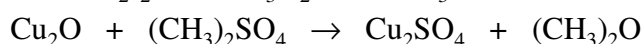
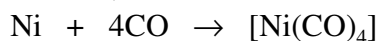
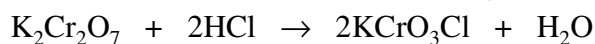
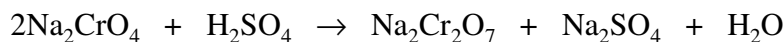
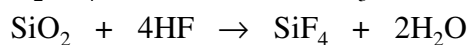
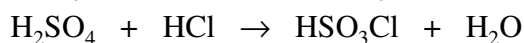
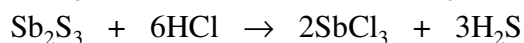
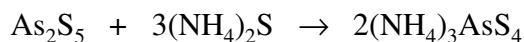
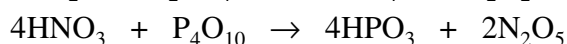
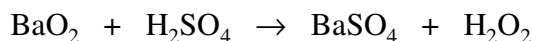
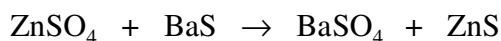
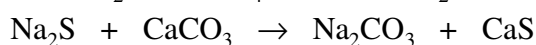
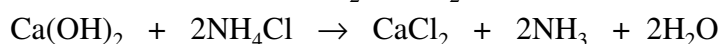
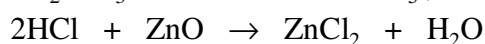
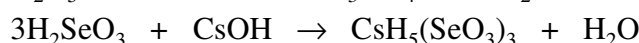
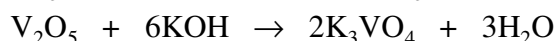
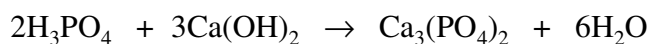
termického rozkladu imidu zinečnatého

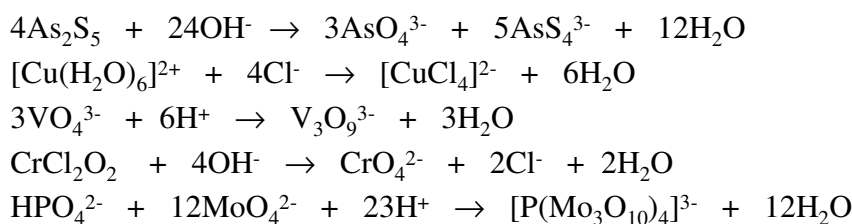


termického rozkladu monohydrátu dihydrogenfosforečnanu sodného



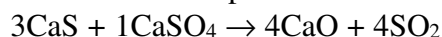
Příklad 2.2.10. Doplňte koeficienty v následujících rovnicích:



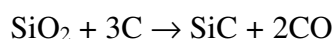


Příklad 2.3.19. Sestavte rovnice následujících oxidačně-redukčních dějů:

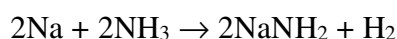
reakce sulfidu vápenatého se síranem vápenatým



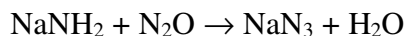
reakce oxidu křemičitého s uhlíkem



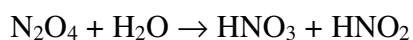
reakce sodíku s amoniakem



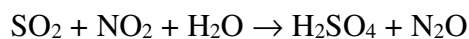
reakce amidu sodného s oxidem dusným



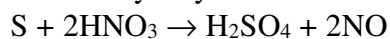
reakce dimerního oxidu dusičitého s vodou



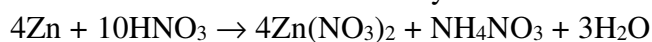
reakce oxidu siřičitého s oxidem dusičitým ve vodném prostředí



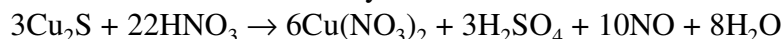
reakce síry s kyselinou dusičnou



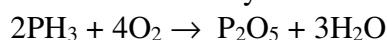
reakce zinku s koncentrovanou kyselinou dusičnou



reakce sulfidu měďného s kyselinou dusičnou



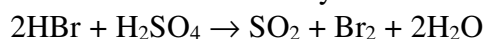
reakce fosfanu s kyslíkem



reakce mědi s koncentrovanou kyselinou sírovou



reakce bromovodíku s kyselinou sírovou



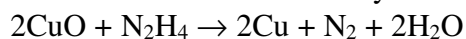
reakce manganistanu draselného s kyselinou chlorovodíkovou



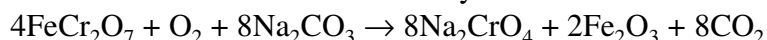
reakce síranu železnatého s peroxidem vodíku v prostředí kyseliny sírové



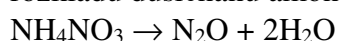
reakce oxidu měďnatého s hydrazinem



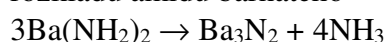
reakce dichromanu železnatého s kyslíkem a uhličitanem sodným



rozkladu dusičnanu amonného



rozkladu amidu barnatého



Příklad 2.3.20. Doplňte koeficienty v následujících oxidačně-redukčních rovnicích:

